

**Average Atomic Mass and Molar Mass**

1. Rubidium has two common isotopes,  $^{85}\text{Rb}$  and  $^{87}\text{Rb}$ . If the abundance of  $^{85}\text{Rb}$  is 72.2% and the abundance of  $^{87}\text{Rb}$  is 27.8%, what is the average atomic mass of rubidium?
2. Uranium has three common isotopes. If the abundance of  $^{234}\text{U}$  is 0.01%, the abundance of  $^{235}\text{U}$  is 0.71%, and the abundance of  $^{238}\text{U}$  is 99.28%, what is the average atomic mass of uranium?
3. Titanium has five common isotopes:  $^{46}\text{Ti}$  (8.0%),  $^{47}\text{Ti}$  (7.8%),  $^{48}\text{Ti}$  (73.4%),  $^{49}\text{Ti}$  (5.5%),  $^{50}\text{Ti}$  (5.3%). What is the average atomic mass of titanium?

Find the molar masses of the following compounds.

4.  $\text{NaBr}$
5.  $\text{PbSO}_4$
6.  $\text{Ca(OH)}_2$
7.  $\text{Na}_3\text{PO}_4$
8.  $(\text{NH}_4)_2\text{CO}_3$
9.  $\text{C}_6\text{H}_{12}\text{O}_6$